

Competitive sorption of molybdate and phosphate in Andisols

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Abstract

We have examined the effects of both pH and phosphate concentrations on molybdate sorption by four Chilean Andisols with varying chemical properties. For both anions adsorption decreased with an increase in soil pH and also with organic matter content in the soils. In general, more phosphate than molybdate was sorbed by the soils. The Freundlich equation adequately described adsorption of both anions. Differences in the extent of sorption of molybdate and phosphate and in isotherm shape between soils were ascribed to variations in soil properties. The competitive sorption data were adequately described by the multicomponent Freundlich equations. Phosphate strongly competed with molybdate for sorption sites on the Andisols, causing molybdate sorption to decrease by 10–27%. The competition coefficients for phosphate and molybdate sorption were variable, indicating that competition between the two anions was not symmetrical. The present results provide further evidence to indicate that the surfaces of Andisols are rich in reactive Al-OH and Fe-OH groups, exposed at colloidal particle surfaces. Molybdate sorption appeared to be mainly associated with free Al- and Fe-oxides, while phosphate sorption was primarily regulated by Al- and Fe-humus complexes.

Keywords: Andisol, Competitive sorption, Molybdate, Phosphate, pH

1. Introduction

The majority of Chilean forestry, agriculture, and livestock are developed on soils derived from volcanic ash, occupying about 2264 km². Andisols account for 65% of the country's land surface, and differ from other mineral soils in their physical, chemical and biological properties which, in turn, determine their agronomical use and management. The mineralogical composition of Andisols is dominated by poorly crystalline minerals such as allophane with variable (pH-dependent) charge. From an agricultural standpoint, these soils have many favorable characteristics, including good physical structure, high organic matter content and water holding capacity, and low bulk density. However the high P retention and soluble Al content may affect agriculture production (Mora *et al.* 2006).

In order to determine the anion nutrient availability in Andisols we studied the anion sorption (Mora *et al.* 2005), and competition between selenite and phosphate (Barrow *et al.*, 2005), and sulphate and phosphate (Jara *et al.* 2006) adsorption. According previous studies (Vistoso *et al.*, 2009; Lopez *et al.*, 2007), P strongly compete with Mo for active sites, thereby increasing Mo availability in Andisols. Molybdenum (Mo) is an essential micronutrient for plant growth, although in high concentrations it can be toxic for livestock. Molybdate ion (MoO_4^{2-}) is the major Mo species in soil solutions and the principal form of Mo that plant roots absorb.

The high anion retention capacity of Andisols affects the mobility and bioavailability of anions such as phosphate and molybdate, thereby resulting in the low efficiency of phosphate and molybdate fertilizers in these soils (Ribera *et al.*, 2010). The relationship between different soil properties and phosphate sorption capacity of Andisols has been extensively investigated (Bolan & Barrow, 1984; Barrow, 1985; Parfitt, 1989) but less information with respect to molybdate sorption. Our preliminary results (Vistoso, 2005; Vistoso *et al.*, 2009)

suggest that the mineralogical composition and organic matter content of the Andisols control the kinetics of the sorption of both molybdate and phosphate. Molybdate appeared to have a high affinity for Fe- and Al-oxides, while phosphate was largely sorbed to Fe- and Al-humus complexes and clay minerals (allophane and ferrihydrite).

The relationship between phosphate and molybdate bioavailability in Andisols was reported by Lopez *et al.* (2007) and Ribera *et al.* (2010). Although phosphate is known to be a more effective competitor than molybdate (Theng, 1971; Barrow, 1992), no models describing competitive sorption between phosphate and molybdate have so far been developed.

Here we use the multicomponent Freundlich equations modified by Barrow *et al.* (2005), based on competitive sorption of binary solute mixtures, to describe molybdate-phosphate competition in Chilean Andisols. Sorption of molybdate is treated as analogous to that of phosphate with both anions competing for the same sorption sites. To this end, we examined: i) the effect of pH on phosphate and molybdate sorption by some Andisols from Southern Chile, and ii) the effect of phosphate on molybdate sorption by Andisols varying in anion sorption capacity.

2. Materials and methods

2.1. Soil collection

Surface samples (0–20 cm depth) of four Andisols (Vilcún, Entic Dystrandep; Pemehue, Hydric Dystrandep; Osorno, Typic Dystrandep and Piedras Negras, Typic Dystrandep) were collected between 37° 35'–44° 04' latitude south and 70° 50'–71° 35' longitude west in Southern Chile. The soils were selected for their contrasting chemical properties and Mo availability to plants. Bulk samples were air-dried, crushed and passed through a 2 mm sieve.

2.2. Physical-chemical characterization of soils

The chemical characteristics of the soil samples were determined according to the methodology modified by Sadzawka *et al.* (2006). Soil pH was measured potentiometrically in water using soil suspensions at a soil:solution ratio of 1:2.5. Organic matter was estimated by wet digestion using a modified Walkley-Black method. Exchangeable base cations (Ca, Mg, K, Na) and exchangeable aluminium were extracted with 1 M NH_4 -acetate at pH 7.0 and 1 M KCl, respectively, and analyzed using atomic absorption spectrophotometry (AAS). Phosphate and sulphate were extracted with 0.5

M NaHCO_3 at pH 8.5 and 1 M $\text{Ca}(\text{H}_2\text{PO}_4)_2$, respectively. Exchangeable Mo was extracted with ammonium bicarbonate diethylene triamine pentacetic acid (AB-DTPA) (Soltanpour, 1991), while total Mo was extracted with HNO_3 and H_2O_2 (Edgell, 1988). Both exchangeable and total Mo were analyzed using graphite furnace atomic absorption spectrometry (GFAAS). The total specific surface area (SSA) was measured by sorption of ethylene glycol monoethyl ether (EGME) (Heilman *et al.*, 1965), while the number of active sites (N_s) for anion sorption was measured by distillation with 1 M Ca-acetate (Holtzclaw & Sposito, 1979). The physico-chemical characteristics of the soils are listed in Table 1.

Table 1. Physico-chemical characteristics of soils.

Parameters	Vilcún	Pemehue	Osorno	P. Negras
Soil classification ^a	Entic Dystrandept	Hydric Dystrandept	Typic Dystrandept	Typic Dystrandept
SSA (m^2/g) ^b	161.59 \pm 1.84	166.10 \pm 1.65	149.95 \pm 1.50	120.26 \pm 1.72
N_s ($\text{cmol}(-)\text{kg}^{-1}$) ^c	30.45 \pm 0.75	36.62 \pm 0.94	70.00 \pm 0.57	58.63 \pm 0.82
pH water ^d	5.47 \pm 0.01	5.41 \pm 0.06	4.81 \pm 0.02	5.40 \pm 0.02
Organic matter (g kg^{-1}) ^d	140 \pm 1.26	150 \pm 0.58	210 \pm 1.16	240 \pm 1.41
S (mg kg^{-1}) ^d	10 \pm 0.64	7 \pm 0.42	20 \pm 0.58	10 \pm 0.25
P (mg kg^{-1}) ^d	13 \pm 0.76	6 \pm 0.62	24 \pm 0.30	5 \pm 0.15
Al ($\text{cmol}(+)\text{kg}^{-1}$) ^d	0.24 \pm 0.01	0.22 \pm 0.05	0.98 \pm 0.02	0.30 \pm 0.04
Σ cations ($\text{cmol}(+)\text{kg}^{-1}$) ^d	8.89 \pm 0.11	4.89 \pm 0.13	4.30 \pm 0.18	3.04 \pm 0.16
Al saturation (%) ^d	2.63 \pm 0.17	4.31 \pm 0.18	18.56 \pm 0.84	8.98 \pm 0.13
Extractable Al (mg kg^{-1}) ^d	600 \pm 11.67	1100 \pm 15.44	1473 \pm 35.69	1876 \pm 16.74
Exchangeable Mo (mg kg^{-1}) ^e	0.009 \pm 0.001	0.010 \pm 0.001	0.016 \pm 0.002	0.022 \pm 0.003
Total Mo (mg kg^{-1}) ^f	0.22 \pm 0.01	0.48 \pm 0.03	0.56 \pm 0.04	0.76 \pm 0.05

^aSpecific surface area (SSA) determined by sorption of ethylene glycol monoethyl ether (Heilman *et al.*, 1965).

^bNumber of active sites (N_s) determined according to methodology described by Holtzclaw and Sposito (1979). ^cDetermined according to methodology described by Sadzawka *et al.* (2006). ^dExtracted with HNO_3 and H_2O_2 (Edgell, 1988). ^eExtracted with ammonium bi-carbonate diethylene triamine pentacetic acid, AB-DTPA (Soltanpour, 1991).

2.3. Molybdate and phosphate sorption isotherms

The sorption isotherms for phosphate and molybdate were obtained from batch experiments using a soil:solution ratio of 1:20 with 0.1 M KCl as back-

ground electrolyte. Single species sorption was measured at initial pH values of 4.0, 5.0 and 6.0, adjusted by adding dilute HCl or KOH solution to the soil suspensions in an N_2 environment so as to eliminate CO_2 . The initial molybdate ($\text{Na}_2\text{MoO}_4 \cdot 2\text{H}_2\text{O}$) and

phosphate (K_2HPO_4) concentrations varied between 0 and 5.0 mmol L⁻¹.

Duplicate samples were weighed and placed in polypropylene centrifuge tubes, and the suspensions were shaken in an end-over-end shaker for 24 h at 298 K. The pH of the suspensions was measured immediately thereafter. The soil suspensions were then centrifuged at 10 000 g for 10 minutes and aliquots of the supernatants were taken for analysis. Mo was determined by AAS and P by the phosphomolybdate-ascorbic acid method (Sadzawka *et al.* 2006). The amount of anion sorbed was derived from the difference between the initial and final solution concentrations.

2.4. Molybdate and phosphate competitive sorption isotherms

The competitive sorption isotherms for phosphate and molybdate were obtained in batch systems using a soil:solution ratio of 1:20 with 0.1 M KCl as background electrolyte at an initial pH of 5.0. For the competitive sorption studies the initial molybdate ($Na_2MoO_4 \cdot 2H_2O$) concentrations were: 0, 0.25, 0.50, 0.75, 1.0, 1.5, 2.0, 2.5, 3.0, 3.5, 4.0, 4.5 and 5.0 mmol L⁻¹, and each of these concentrations were tested with 0, 0.5, 1.0 and 2.0 mmol L⁻¹ phosphate (K_2HPO_4). Duplicate samples were weighed and placed in polypropylene centrifuge tubes, and the suspensions were shaken in an end-over-end shaker for 24 h at 298 K. Suspension pH and the concentrations of Mo and P in the supernatant solutions were determined as described above.

2.5. Mathematical sorption modelling

Single species molybdate and phosphate sorption data at different soil pH levels ($n = 4$) were fitted to the Freundlich equation (Eq. 1):

$$X = a C^{(1/b)} \quad (1)$$

where X is the amount of anion sorbed (mmol kg⁻¹); C is the equilibrium anion concentration (mmol L⁻¹), and a and b are constants.

The competitive sorption data for Mo and P for each of the four soils were fitted to the multicomponent Freundlich equations (Eqs. 2 and 3) using a computer program, Model 2 modified for the competitive sorption of Mo and P, (Barrow *et al.*, 2005):

$$S_{Mo} = a_{Mo} C_{Mo} / (C_{Mo} + K_{Mo,P} C_P)^{1-b_{Mo}} \quad (2)$$

$$S_P = a_P C_P / (C_P + K_{P,Mo} C_{Mo})^{1-b_P} \quad (3)$$

where S_{Mo} is the amount of Mo sorbed per unit mass of sorbent in the presence of competitor P, and S_P is the amount of P sorbed per unit mass of sorbent in the presence of competitor Mo; a_{Mo} and a_P are the single-solute Freundlich constants for solutes Mo and P; C_{Mo} and C_P are the equilibrium concentrations – ideally activities – of the solutes; $K_{Mo,P}$ is the competition coefficient of P on Mo; and $K_{P,Mo}$ is the competition coefficient of Mo on P. According to Barrow *et al.* (2005), molybdate-phosphate competition was not symmetrical. Furthermore, the process involved not only competition for sorption sites but also electrical potential effects following penetration of the adsorbent's structure by the solute.

Sorption was calculated from the observed change in solution concentration. Sorption and solution concentration are therefore not independent variables. This problem was circumvented by regarding the above sorption equations as simultaneous with the equation used to calculate sorption from the observed change in concentration. For any set of parameters, a simultaneous result was found by iteration using the computer program. This allowed the final equilibrium concentration to be predicted for each initial concentration. A sim-

ple routine was then used to find the sets of parameters that minimized the residual sum of squares of the difference between observed and predicted equilibrium concentrations.

2.6. Statistical analysis

The experimental data were statistically analyzed using SPSS version 13. The fit of each mathematical sorption model was evaluated by determining the coefficient (R^2) and standard error (SE) and used to compare the applicability of the different models at a confidence level of 95%.

3. Results and Discussion

3.1 Effect of pH on molybdate and phosphate sorption

The sorption isotherms for both anions as a function of pH are presented in Figures 1 and 2. For both anions, sorption decreased with an increase of pH in suspension. At any given pH and solution concentration, more phosphate than molybdate was sorbed by all Andisols except for the Vilcún soil.

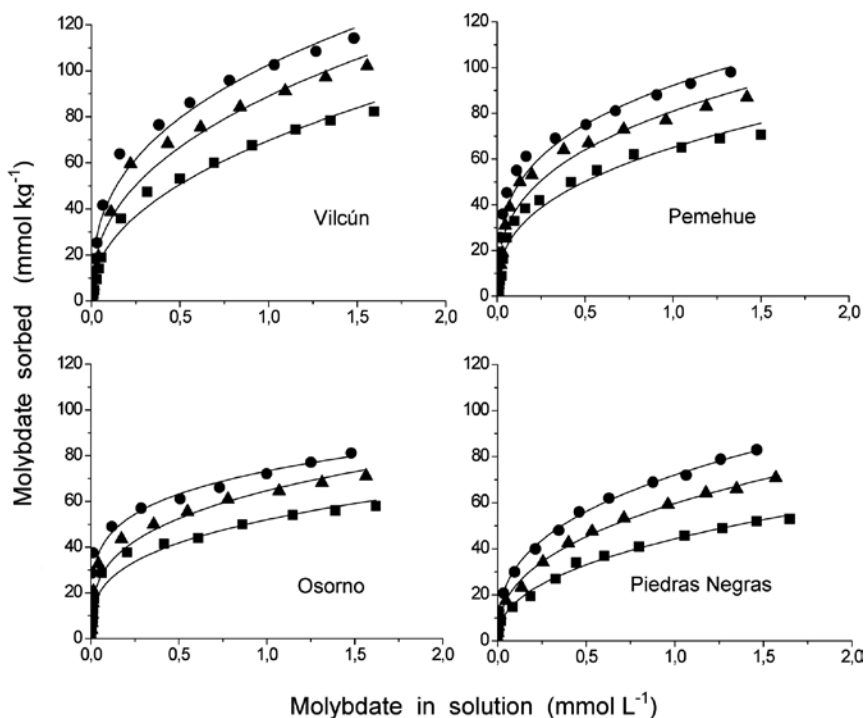


Figure 1. Effect of initial pH on molybdate sorption by Vilcún, Pemehue, Osorno and Piedras Negras soils (● pH 4.0; ▲ pH 5.0; ■ pH 6.0: Experimental data; —: Model 2 fitted lines).

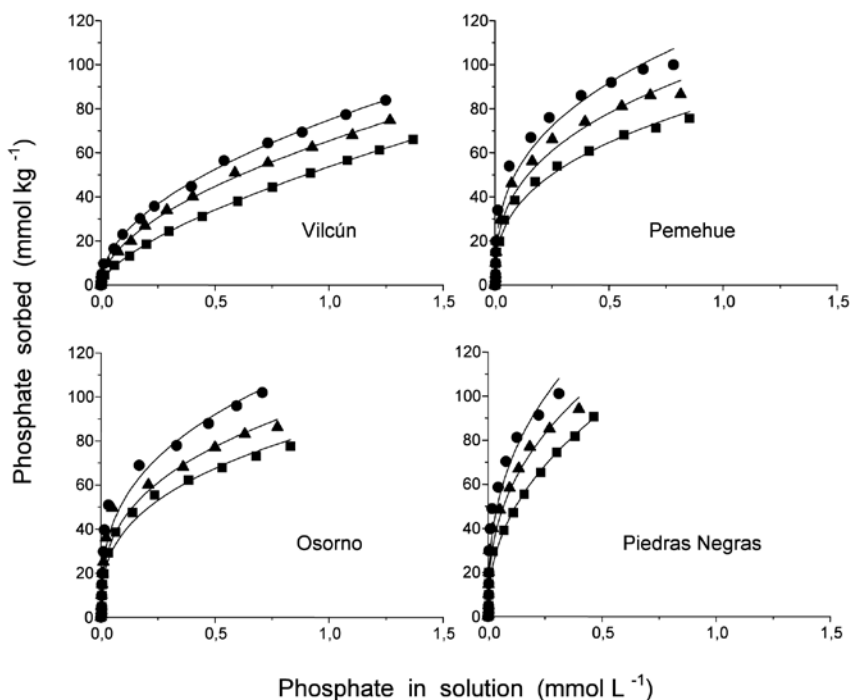


Figure 2. Effect of initial pH on phosphate sorption by Vilcún, Pemehue, Osorno and Piedras Negras soils (● pH 4.0; ▲ pH 5.0; ■ pH 6.0: Experimental data; —: Model 2 fitted lines).

Taking sorption at pH 4.0 as 100%, molybdate sorption decreased by 12% at pH 5.0 and 26% at pH 6.0, while phosphate sorption decreased by 16% at pH 5.0 and 32% at pH 6.0. Similarly, Goldberg *et al.* (2008) and Goldberg *et al.* (2009) noted a marked decrease in P and Mo sorption by oxides and volcanic soils at $\text{pH} > 5.0$, ascribed to differences in the origin, mineral composition, and degree of development of the soils.

For both anions, the initial slope of the isotherms depends on solute concentration. The curves also indicate a high anion-surface affinity. The isotherms for molybdate sorption tend to be of the L-type (Giles *et al.*, 1974), indicating that the number of high affinity sites diminishes as sorption progresses. Less molybdate was sorbed by the Vilcún and Pemehue soils

than by the Osorno and P. Negras soils, consistent with the lower number of active sites (N_s) determined for the former two soils (Table 1). On the other hand, the phosphate sorption isotherms tend to be of the H-type, indicating high anion-surface affinity especially for the Osorno and P. Negras soils. These two soils also have a relatively large number of active sites although their specific surface areas (SSA) are smaller than those of the Vilcún and Pemehue soils.

For all the Andisols used except Vilcún, the extent of molybdate sorption was 20% lower than that of phosphate (Figures 1 and 2). Molybdate sorption was accompanied by liberation of hydroxyl ions into the soil solution, increasing the supernatant pH by up to one unit. The decrease in molybdate sorption in the

presence of phosphate may be attributed to the two factors that include steric hindrance and electrostatic repulsion; these factors might be expected to cause the difference in sorption behaviour of molybdate by these Andisols, relative to phosphate.

Barrow (1970) found that the depressing effect of increased pH on molybdate sorption was about 20 times greater than for phosphate. The difference between our and Barrow's (1970) observation may arise because we used 0.1 M KCl, whereas Barrow (1970) used 0.01 M CaCl_2 . The decrease in electric potential with increasing pH is less in a calcium system than in a potassium one. On the other hand, Golberg *et al.* (2009) reported that Mo sorption from pH 4.0 to 8.0 by a range of soils was independent ionic strength.

The observed increase in supernatant pH due to Mo sorption (Figure 3a) clearly indicates that the underlying mechanism is one of ligand exchange between molybdate ions in solution and surface hydroxyl groups, similar to that proposed for phosphate (Goldberg & Sposito, 1985; Parfitt, 1989; Goldberg *et al.*, 2008; Vistoso, 2005; Vistoso *et al.*, 2009). As a result, the net negative surface charge would increase (Bolan & Barrow, 1984; Bolan *et al.*, 1986; Theng *et al.*, 2005), while the sorption capacity of the soils for both P and Mo anions would decrease (Figures 1 and 2). Similar behaviour is observed with increasing supernatant pH due to competitive sorption of molybdate in the presence of phosphate for all the Andisols (Figure 3b), in agree with Goldberg & Forster (1998).

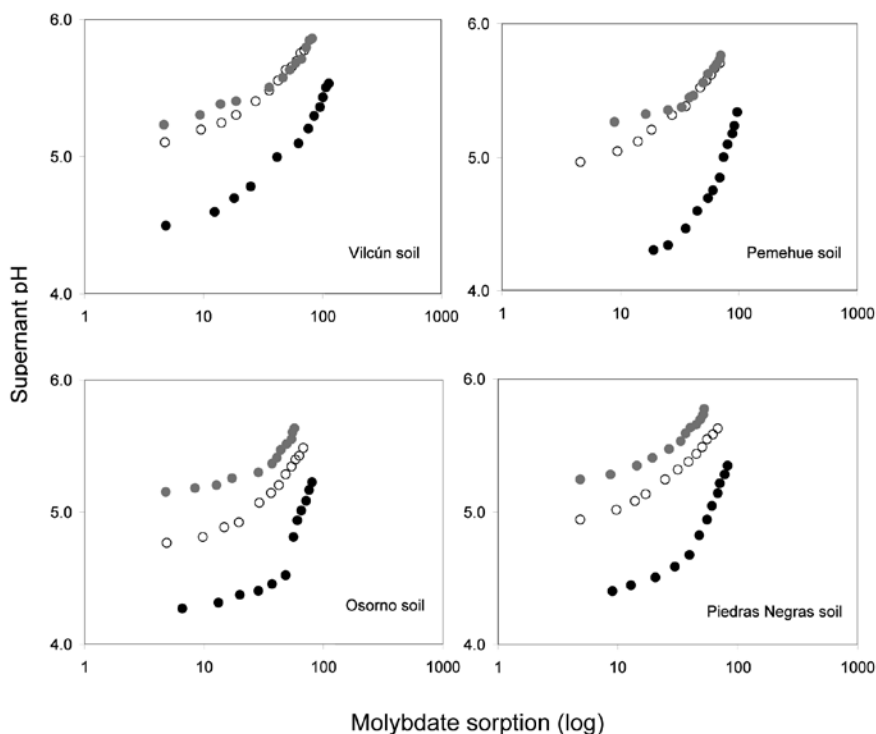


Figure 3a. Relationship between supernatant pH and molybdate sorbed by Vilcún, Pemehue, Osorno and Piedras Negras soils at different initial pH levels (● pH 4.0; ○ pH 5.0; ● pH 6.0).

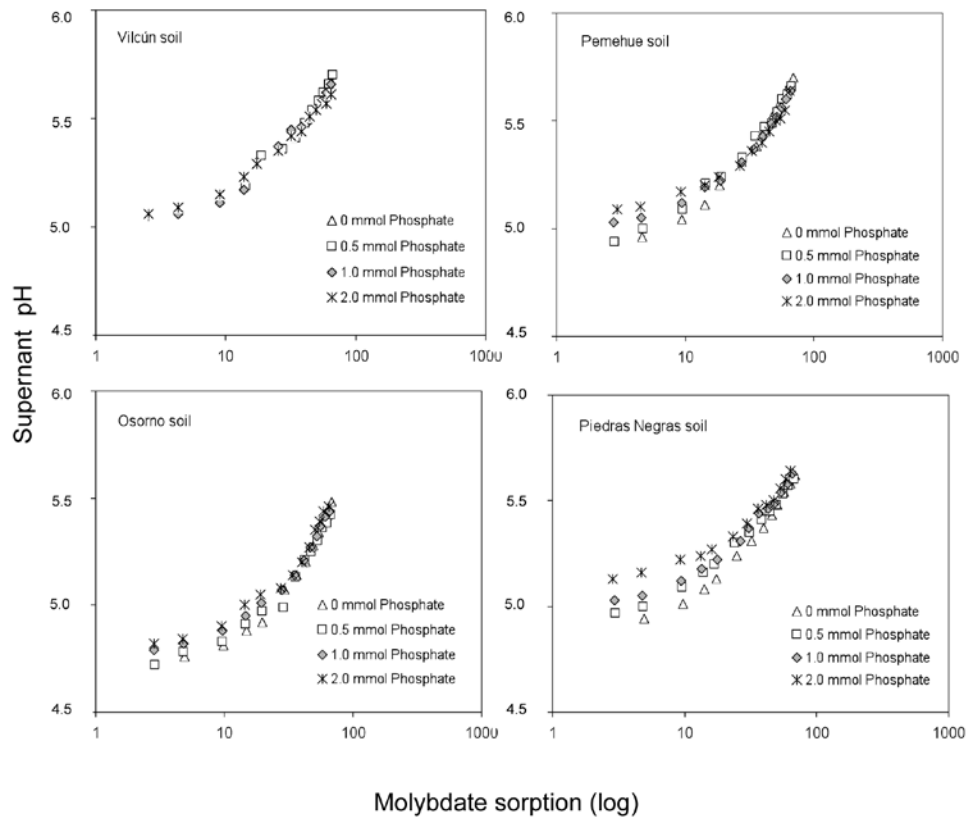


Figure 3b. Relationship between supernatant pH and molybdate sorbed in presence of different concentration of phosphate by Vilcún, Pemehue, Osorno and Piedras Negras soils at initial pH 5.0.

The experimental data for both anions were adequately modelled using the Freundlich equation (Figures 1 and 2) with $R^2 > 0.921$ (Table 2). For the Andisols used, the sorption parameters indicate that both the sorption

capacity (*a*), and sorption energy (*b*) for phosphate are greater than that for molybdate except for the Vilcún soil. Both parameters decreased in magnitude as sorption declined with rising pH (Table 2).

Table 2. Freundlich isotherm parameters for molybdate and phosphate sorption at different initial pH values.

Soil	pH 4.0			pH 5.0			pH 6.0		
	Parameters (mmol kg ⁻¹)								
	<i>a</i>	<i>b</i>	R ²	<i>a</i>	<i>b</i>	R ²	<i>a</i>	<i>b</i>	R ²
Molybdate									
Vilcún	102.56	0.366	0.983	88.95	0.411	0.980	69.66	0.459	0.982
Pemehue	92.26	0.298	0.966	80.94	0.339	0.951	65.03	0.373	0.951
Osorno	73.25	0.213	0.984	64.63	0.299	0.971	52.00	0.313	0.964
P. Negras	72.03	0.359	0.997	59.65	0.399	0.996	44.30	0.424	0.995
Phosphate									
Vilcún	74.87	0.512	0.998	65.45	0.549	0.998	53.67	0.657	0.999
Pemehue	116.70	0.347	0.963	100.01	0.358	0.972	83.45	0.367	0.979
Osorno	116.69	0.342	0.946	97.97	0.341	0.948	85.97	0.348	0.971
P. Negras	167.63	0.376	0.921	142.17	0.388	0.948	125.99	0.438	0.975

These observations are consistent with those reported by Goldberg & Forster (1998) and Barrow (1999) who showed that phosphate had a greater affinity than molybdate for a range of soils varying in Fe- and Al-oxide contents and pH levels. The present results also accord with those derived from kinetic measurements, using the same anions and soils (Vistoso *et al.*, 2009).

In conclusion, the reactivity of molybdate toward the Andisols decreased in the following order: Vilcún > Pemehue > Osorno > Piedras Negras. On the other hand, the reactivity of phosphate decreased in the opposite order: Piedras Negras > Osorno > Pemehue > Vilcún. These results corroborate previous findings by Vistoso *et al.* (2009) that molybdate ion sorption affinity in the Vilcún and Pemehue soils was related to the content of free (dithionite-extractable) Al- and Fe-oxides (23.00 and 40.10 g kg⁻¹ of Al_d + Fe_d, respectively). These workers further showed that phosphate availability in the Osorno and Piedras Negras soils was principally regulated by pyrophosphate-extract-

able Al- and Fe-humus complexes (23.00 and 25.00 g kg⁻¹ of Al_p + Fe_p, respectively), allophane and ferrihydrite. Likewise, Golberg *et al.* (2008) have suggested that Al- and Fe- oxides, clay minerals, organic matter, and allophane play an important role in molybdate sorption by mineral and volcanic ash soils.

3.2. Effect of phosphate on molybdate sorption

The analysis of competitive sorption indicated that molybdate sorption decreased by 10–27% in the presence of phosphate (0.5–2.0 mmol L⁻¹) for all the Andisols (Figure 4). The reduction in molybdate sorption was greater at higher levels of phosphate addition, suggesting strong competition between the two anions. Previously, Lopez *et al.* (2007) showed in incubation studied the Mo availability increased by phosphate addition in Andisols. The phosphate-induced decrease in molybdate sorption was higher for the Pemehue and Piedras Negras soils than the Vilcún and Osorno soils (Figure 4).

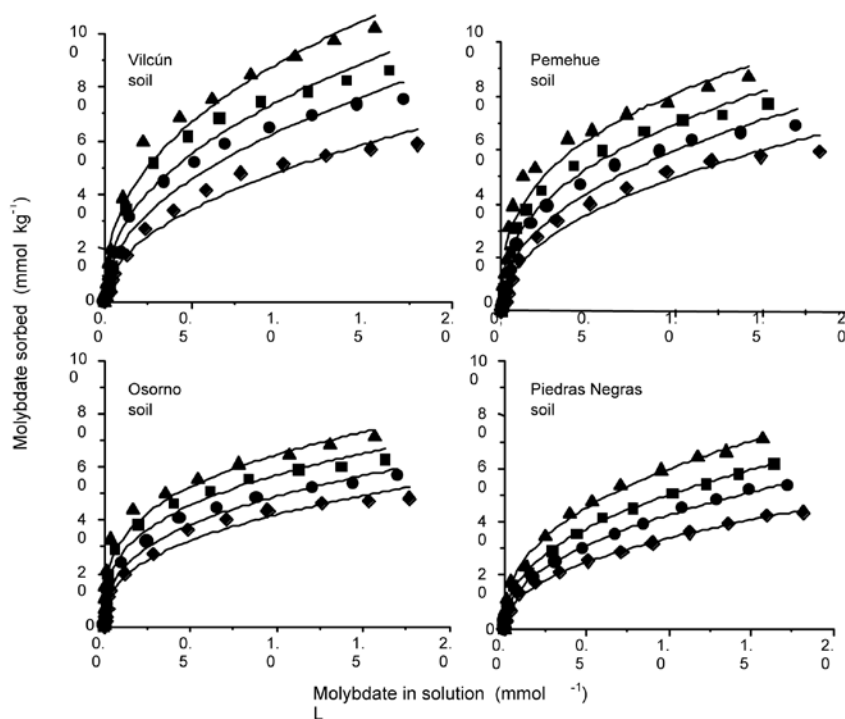


Figure 4. Competitive sorption of molybdate in presence of different concentrations of phosphate by Vilcún, Pemehue, Osorno and Piedras Negras soils (—:Model 2 modified lines; Experimental data: P1 (▲) 0 mM P, P2 (■) 0.5 mM P, P3 (●) 1.0 mM P, P4 (◆) 2.0 mM P).

Competitive interactions between phosphate and molybdate indicated that at low molar ratios molybdate and phosphate compete for the same set of surface sites as reflected in a lower $K_{P,Mo}$ (competition of Mo on P) and higher $K_{Mo,P}$ (competition of P on Mo) coefficients. Further, the reduction in molybdate sorption was greater at higher levels of phosphate additions, suggesting strong competition in the Andisols used. Similarly, Xie & McKenzie (1991) and Xu *et al.* (2006) observed that molybdate sorption to soils, goethite and pyrite was significantly reduced by the competitive sorption of phosphate.

At high concentrations phosphate was more effective in inhibiting molybdate sorption. This might

be because sorption of phosphate by ligand exchange increased the net negative surface charge, enhancing electrostatic repulsion between molybdate and colloidal particle surface (Barrow, 1999; Barrow *et al.*, 2005). These results are therefore contrary to those obtained by Roy *et al.* (1986) who suggested that phosphate and molybdate would continue to compete for the same surface sites at high phosphate concentrations.

The close fit of the competition data to the Freundlich equation, indicated by the R^2 values of 0.947–0.996 for both molybdate and phosphate (Table 2), suggests the model can adequately describe molybdate sorption in the presence of added phos-

phate. At the same time, the $K_{Mo,P}$ and $K_{P,Mo}$ values are indicative of non-symmetrical competition between Mo and P, rather than simple competition for sorption sites.

The ability of phosphate to out compete molybdate in these Andisols decreased in the order:

Pemehue > Piedras Negras > Vilcún > Osorno. The coefficient values (Table 3) and Figure 5 show that phosphate sorption by the Andisols was not affected by different molybdate concentrations. By contrast, molybdate sorption was sensitive to the presence of phosphate (Figure 6 and Table 3).

Table 3. Parameters of the multicomponent Freundlich model fitted for competitive sorption between molybdate and phosphate in Andisols.

Soil	Parameters				$K_{Mo,P}$	$K_{P,Mo}$	R^2_{Mo}	R^2_P
	a_{Mo}	a_P	b_{Mo}	b_P				
Vilcún	57.56	61.49	0.36	0.31	0.01	3.30	0.991	0.995
Pemehue	60.69	82.78	0.31	0.43	0.03	6.03	0.947	0.994
Osorno	57.43	85.34	0.29	0.44	0	2.85	0.996	0.995
P. Negras	51.96	134.54	0.50	0.49	0.01	4.53	0.996	0.990

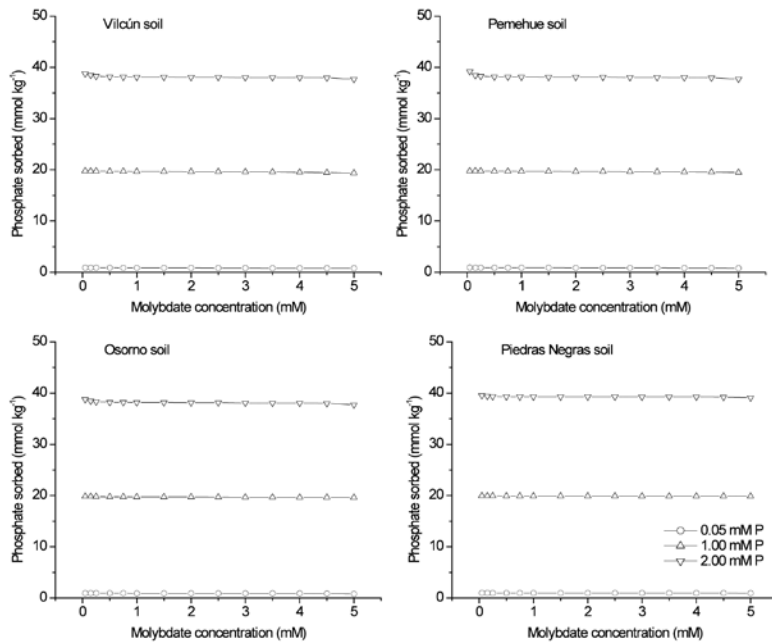


Figure 5. Effect of different molybdate concentrations on phosphate sorption by Vilcún, Pemehue, Osorno and Piedras Negras soils (○: 0.05 mM P; △: 1.00 mM P; ▽: 2.00 mM P).

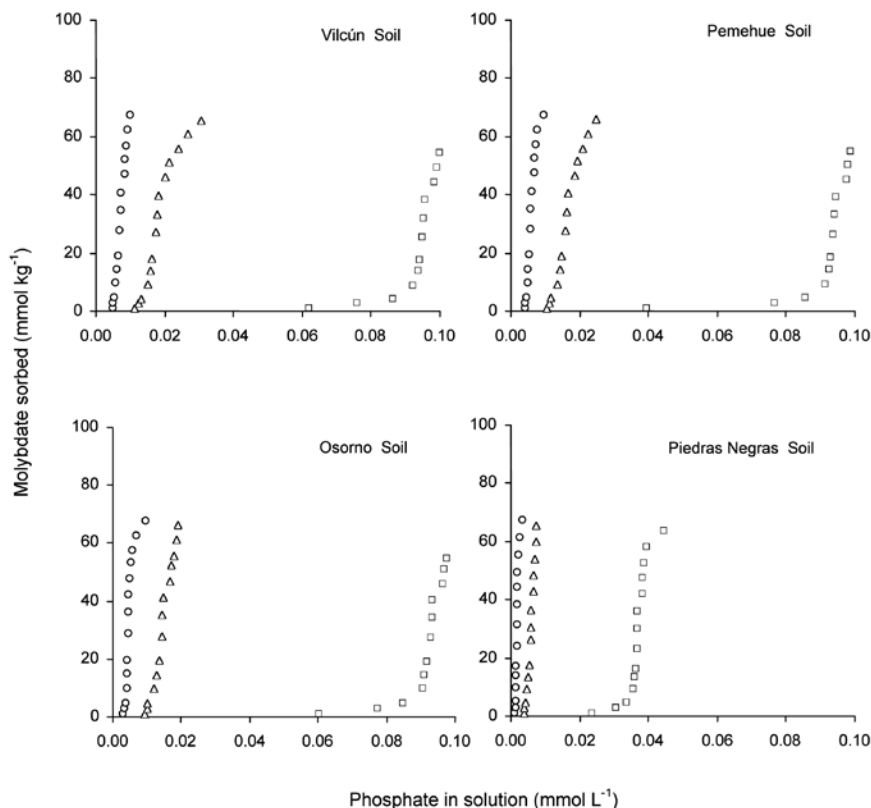


Figure 6. Effect of different phosphate concentrations on molybdate sorption by Vilcún, Pemehue, Osorno and Piedras Negras soils (\circ : 0.05 mM P; Δ : 1.00 mM P; \square : 2.00 mM P).

The superior ability of phosphate to compete against molybdate in the Andisols used may be explained in following terms: a) differences in the capacity of some soil components for sorbing phosphate and molybdate; b) since phosphate and molybdate are sorbed to similar surface sites and through a similar mechanism, increasing net negative surface charge, enhancing electrostatic repulsion between molybdate and colloidal particle; c) phosphate can access more surface sites than molybdate; d) phosphate and molybdate combined may be sorbed in greater amounts than each anion separately.

Conclusions

The sorption of molybdate and phosphate by some Chilean Andisols is sensitive to solution pH and may be modeled by the Freundlich equation. In general, much less molybdate than phosphate was sorbed by the soils. The capacity of the soils for sorbing molybdate and phosphate is related to their chemical properties. The increase in solution pH, accompanying molybdate sorption, indicates that this anion are sorbed through ligand exchange with hydroxyl groups exposed at colloidal particle surfaces of the Andisols.

Phosphate competes strongly with molybdate for sorption sites resulting in more molybdate concentration in soil solution, increasing Mo availability for plant production. Molybdenum deficiency in Andisols may therefore be overcome by maintaining its concentration in soil solution at adequate levels through phosphate fertilization.

Acknowledgements

This research was supported by the Comisión Nacional de Investigación Científica y Tecnológica, CONICYT and Fondo Nacional de Desarrollo Científico y Tecnológico, FONDECYT through the projects N° 1100625 and MEC-CONICYT N° 80100008.

We are grateful to Dr. N.J. Barrow of Mt. Claremont, Australia for assistance in modelling competitive anion sorption by soils.

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