

DISCLAIMER

This document was prepared as an account of work sponsored by the United States Government. While this document is believed to contain correct information, neither the United States Government nor any agency thereof, nor The Regents of the University of California, nor any of their employees, makes any warranty, express or implied, or assumes any legal responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by its trade name, trademark, manufacturer, or otherwise, does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof, or The Regents of the University of California. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof or The Regents of the University of California.

Self-assembly of uniform polyhedral silver nanocrystals into densest packings and exotic superlattices

Joel Henzie,¹ Michael Grünwald,^{1*} Asaph Widmer-Cooper,^{1,3,4}
Phillip L. Geissler,^{1,3} Peidong Yang^{1,2,3*}

¹Department of Chemistry, University of California,
Berkeley, California 94720, USA.

²Department of Materials Science and Engineering, University of California,
Berkeley, California 94720, USA.

³Materials Sciences Division, Lawrence Berkeley National Laboratory,
Berkeley, California 94720, USA.

⁴School of Chemistry, University of Sydney,
Sydney, New South Wales 2006, Australia.

*Corresponding authors;

E-mail: p.yang@berkeley.edu (P.Y.), michael.gruenwald@berkeley.edu (M.G.)

Understanding how polyhedra pack into extended arrangements is integral to the design and discovery of crystalline materials at all length scales [1, 2, 3]. Much progress has recently been made in enumerating and characterizing the densest packings of polyhedral shapes [4, 5, 6]. While polyhedral nanocrystals have been assembled into ordered superstructures [7, 8, 9], directing their self-assembly into densest packings requires precise control of particle shape [10], polydispersity [11], interactions, and driving forces [12]. Here we show with experiment and computer simulation that a range of nanoscale Ag polyhedra can self-assemble into their conjectured densest packings [6]. When passivated with adsorbing polymer, they behave as quasi-hard particles, and assemble into millimeter sized 3D supercrystals by sedimentation. By inducing depletion attractions through excess polymer in solution, we can drive octahedra to form an exotic superstructure with complex helical motifs rather than the densest Minkowski lattice [13]. Such large-scale Ag supercrystals provide many possibilities for designing scalable 3D plasmonic metamaterials with applications in sensing [14, 15], nanophotonics [16], and photocatalysis [17].

New synthetic methods give access to a wide range of well-defined polyhedral nanocrystalline shapes. The polyol synthesis was used to generate monodisperse Ag nanocrystals (Supplementary Methods) including cubes, truncated cubes, cuboctahedra, truncated octahedra, and octahedra over a range of sizes from 100–300 nm (Fig. 1a) [18, 7]. These Ag nanocrystals are coated with polyvinyl-pyrrolidone (PVP; 55,000 MW), which enables the shape-controlled synthesis and acts as a stabilizer [19]. Of these five shapes, only cubes and truncated octahedra possess space-filling packings. For the other three shapes, as for most non-spherical shapes, no rigorous statements about optimal packings exist.

For any convex shape, however, the *densest lattice packing* can be calculated explicitly [4, 13]. Lattice packings are a special class of arrangements in which all particles are oriented identically and positioned at sites of a crystalline lattice. While it is generally not true that the densest lattice packing constitutes the overall densest packing for a given shape, a recent conjecture by Torquato and Jiao suggests that the centrally symmetric Platonic and Archimedean solids—families of convex polyhedra that include the shapes studied here but, notably, do not include tetrahedra—pack most densely into these lattice packings [6]. Yet this statement does not guarantee that such structures are thermodynamically stable at finite pressure, nor that they can be readily accessed through random thermal motion. Indeed, computer simulations by Haji-Akbari et al. show, in the related case of tetrahedra, that these factors result in self-assembly of an intriguing quasi-crystal that is not the densest packing [5, 20, 21, 22].

In our experiments, monodisperse Ag polyhedra assemble into large ($> 25 \text{ mm}^2$), dense supercrystals by gravitational sedimentation (Fig. 1b). This process can be carried out in bulk solution, but assembly in polydimethylsiloxane reservoirs (Supplementary Methods) enables precise control of the superlattice dimensions and monitoring by dark-field optical microscopy (Fig. 1c). N,N-Dimethylformamide (DMF) was used as the assembly medium because it is an excellent solvent for PVP. In a typical experiment, a dilute solution of particles was loaded into the reservoir so that the channel filled with a uniform solution of particles (Supplementary Methods). Then the device was tilted, causing the particles to gradually sediment and assemble at the bottom of the reservoir (Fig. 1c; Supplementary Movies 1–3); more concentrated solutions or higher angles of tilt cause the assemblies to form more quickly.

The assemblies generated by this sedimentation procedure, shown in Fig. 1 d-h, exhibit translational and rotational order over exceptional length scales, as illustrated in Fig. 2

a-d. In the cases of cubes, truncated octahedra, and octahedra, the structures of the dense supercrystals correspond precisely to their densest lattice packings. The same packings are observed to form in constant pressure Monte Carlo simulations (Supplementary Methods) of dense fluids of hard polyhedra (Supplementary Figures 2, 5; Supplementary Movie 4). Ag cubes (Fig. 1d) form dense arrangements that fill space. In small assemblies, cubes tend to occupy the sites of a simple cubic lattice and form supercubes of up to 10 microns edge length. In larger assemblies the symmetry of the simple cubic lattice is broken and more irregular, space-filling packings are observed. Truncated octahedra assemble into the space-filling Kelvin structure (Fig. 1g), which corresponds to the body centered cubic lattice. Of all the shapes studied here, assemblies of truncated octahedra exhibit single-crystalline order over the largest length scales. Ag octahedra (Fig. 1h) assemble into a packing that was recognized long ago by Minkowski as the densest lattice packing for this shape [13]. Unlike cubes and truncated octahedra, octahedra in their densest lattice packing engage in only incomplete face-to-face contact; equivalent faces of neighboring octahedra also do not lie in a common plane. As discussed later, these features allow simple biases to stabilize alternative packings. In Figure 2e, a rendering of the perfect Minkowski lattice is superimposed on the experimental lattice, showing excellent agreement. A detailed comparison of Fourier patterns is shown in Supplementary Figure 1.

Cuboctahedra and truncated cube Ag nanocrystals also adopt very dense, ordered structures that closely resemble their densest lattice packings (Fig. 1e-f). For truncated cubes, the densest packing is only subtly different from the simple cubic lattice and their packing fractions differ by only 0.7%. We therefore expect that high pressures are necessary to bias the assembly towards the densest packing. In our experiments, truncated cubes form lattices that vary from almost perfect cubic alignment to the characteristic

tilt of the densest packing. Monte Carlo simulations of truncated cubes similarly exhibit variability at moderate, experimentally accessible pressures. In fact, even at very high pressures the densest packing was never observed to form spontaneously from the fluid on the timescale accessible to straightforward simulation. By contrast, simulations of hard cuboctahedra readily achieve the densest lattice packing at moderately high pressures while experimentally assembled lattices vary between the densest packing and a slightly distorted lattice with face-centered-cubic like symmetry (Supplementary Figure 3).

Calculations confirm that gravity is indeed a plausible driving force in our assembly experiments. Unlike smaller nanocrystals used in conventional, evaporation-driven assembly studies, our Ag particles are 100–300 nm in size and experience substantial gravitational forces, for octahedra amounting to $0.23 k_B T$ per micron. To estimate the pressures that develop within the sediment, we performed Monte Carlo simulations of octahedra under the influence of gravity (Supplementary Methods), in an elongated box with a hard bottom wall (Fig. 3a). To mimic the experimental setup, simulations were started at uniform low density, allowing the particles to sediment, eventually achieving a density gradient approximately 150 microns in height. Resulting pressures near the bottom of the model cell were sufficient for spontaneous crystallization into the Minkowski lattice. Experimental sedimentation channels feature even higher columns of dense material, and therefore pressures more than sufficient to drive crystallization. Further evidence comes from experimental waiting times. While the larger octahedra (300 nm edge length, 1.33×10^{-13} g) sediment within a few hours, the smaller cubes (122 nm edge length, 1.90×10^{-14} g) take significantly longer to assemble.

Both in experiment and simulation, octahedra packed against a flat substrate form a structure that is different from but closely related to the Minkowski lattice (Fig. 3c-e). This simple hexagonal lattice has a maximum packing fraction of $8/9$ and is composed of

flat, hexagonally close-packed layers that stack with perfect face-to-face contact between the layers (Fig. 3e). Octahedra in the corresponding layers of the Minkowski lattice (Fig. 3e, bottom right) are slightly rotated which results in increased packing efficiency (18/19) through interpenetration of stacked layers. In the presence of a flat wall, however, local packing efficiency is maximized by flattening the layers. This boundary effect extends only a short distance into the sample. Both in experiment and simulation, the simple hexagonal packing type gives way to the Minkowski structure after a few close-packed layers.

That structures determined by shape and high pressure can be achieved in our self-assembly experiments indicates that particles are behaving as almost purely hard, nonattractive polyhedra. This conclusion is surprising, given the potentially overwhelming van der Waals (vdW) attractions between metallic objects of this size. At nanometer separations, the corresponding potential energy reaches tens of thousands of $k_B T$. If additional repulsive forces did not preclude such separations, experiments would result in rapid aggregation of disordered, irreversibly bound particle aggregates. By contrast, assemblies are quite reversible prior to the drying step: When the assembly experiment is turned upside down, the top of the assembled lattice dissociates and the nanocrystals begin to assemble at the "new" bottom of the setup (Supplementary Figure 8). Significant electrostatic repulsion is unlikely since the measured zeta potential of the particles in DMF is close to zero (Supplementary Methods). Instead, experimental and computational evidence strongly suggests that the repulsion originates in volume exclusion of the adsorbing polymer PVP.

Measurements of the hydrodynamic diameter of cuboctahedra with adsorbed polymer indicate that particles are covered with a PVP layer that extends 20 nm from the particle surface, which is roughly equal to the hydrodynamic diameter of free PVP in DMF

(Supplementary Methods). Our calculations suggest that polymer layers of this thickness are sufficient to prevent close approach of nanocrystal surfaces. Figure 4a shows model predictions of different contributions to the interaction between two Ag octahedra, aligned face-to-face, as a function of their separation (Supplementary Methods, Supplementary Figure 6). The particles are assumed to be covered with a 20 nm thick layer of dense polymer. As the polymer layers of the two octahedra come into contact, strong entropic repulsion competes with the vdW attraction of the Ag cores. The combined potential energy curve shows that the polymer-mediated repulsion is indeed sufficient to balance the vdW forces, resulting in a shallow effective energy minimum upon contact between the surface polymer layers. To test the influence of the size of the adsorbed polymer brush on this force balance, we have repeated assembly experiments with ethanol and water as solvents. Measurements of the hydrodynamic diameter of cuboctahedra suggest brush thicknesses of roughly 15 nm in ethanol and 5 nm in water (Supplementary Methods). Calculations show that upon contact of the polymer brushes of two octahedra, the vdW energies amount to roughly $-14 k_B T$ in ethanol and $-140 k_B T$ in water (compare to $-7 k_B T$ in DMF). The results of assembly experiments in these solvents are consistent with the estimated increasing attraction strengths: While we do observe some assembly of the Minkowski lattice in ethanol, albeit of inferior quality, we observe primarily disordered aggregation in water.

The effect of the residual vdW attractions is visible in many aspects of the assembly experiments: Optical microscopy shows that particles can temporarily bind even at moderate concentrations (Supplementary Movie 1). Moreover, as discussed earlier in the text, cubes and truncated cubes show a preference for face-to-face alignment in small assemblies, as do close-packed octahedra layers near substrates. However, under conditions of high density, these weak, slowly varying attractive forces play a secondary role [23].

The nanocrystals behave essentially as hard particles with polyhedral shapes that are not modified or masked by the thin layer of adsorbed PVP. Nevertheless, the final dried assemblies shown in Figure 1 are quite stable with inter-particle spacings of only a few nanometers. Surface polymer molecules, driven by high pressures and decreasing solvent concentration, evidently collapse or desorb at later stages of the experiments, allowing particles to approach closely and attract strongly (Supplementary Figure 4).

Unadsorbed polymer in solution can also induce attractions between particles [24]. These depletion attractions are entropic forces generic to solutions of differently sized particles: At small separations, two of the larger particles exclude smaller ones from the intervening volume [25, 26, 27]. The resulting gradient in osmotic pressure (due to the gradient of density of the smaller species) induces an attraction between the large particles that increases with the bulk concentration of the smaller species. In our case, free polymer coils in solution act as depleting agents and mediate substantial attractions between polymer-coated nanocrystals: Figure 4b shows effective pair potentials of hard octahedra in a bath of hard spheres, mimicking free polymer coils in solution, for several different volume fractions φ of the spheres (Supplementary Methods). Note that the typical oscillatory behavior of depletion forces is not evident in these potentials because they are averaged over all rotational degrees of freedom and polymer concentrations are low.

Experiments and simulations indicate that these tunable attractions can induce novel structures that are not governed solely by packing efficiency of hard shapes. Figure 5a-b shows SEM micrographs of a lattice of octahedra with intriguing helical motifs that was generated by adding $0.1 \mu\text{g}/\mu\text{L}$ (w/v) of extra PVP to the solution immediately prior to sedimentation. Higher amounts of PVP resulted in less Minkowski lattice and more helical lattice. The large degree of face-to-face contact in the latter reflects significant attractive

forces. The same structure was recovered in Monte Carlo simulations employing the effective depletion pair potentials shown in Fig. 4b, for $\varphi > 0.02$. (Further evidence for the stability of the helical structure in the presence of depleting agents is discussed in the Supplementary Methods and Supplementary Figure 7.) Symmetry analysis of an optimized unit cell revealed that the structure has space group $I\bar{4}3d$, with a cubic unit cell containing 16 octahedra (Fig. 5c-f, Supplementary File 1). Particle centers are located on 16c Wyckoff sites, similar to a high pressure phase of Lithium [28]; vertices occupy 48e sites. The maximal packing fraction of this novel structure ($\sim 82\%$) is significantly lower than that of the Minkowski lattice ($\sim 95\%$). This helical structure may also be more suitable as a plasmonic bandgap crystal because the lines and helices should create continuous transmission pathways with lower attenuation [7].

The approach we have used in this study circumvents several issues that have been a challenge for reproducible and tunable assembly of nanocrystal patterns over device-length scales. In many experiments the evaporation of solvent triggers organization of nanocrystal solutes. Such drying-mediated assembly can be difficult to control due to the severity of solvation forces and their spatially non-uniform fluctuations [29]. In fact, we have tried several experimental geometries where solvent evaporation is a major driving force. In the simplest examples, a drop of nanocrystals in DMF was mounted on a substrate and evaporated in air, or placed in a atmosphere saturated with DMF vapor to slow evaporation. While we observe the same types of crystals in these experiments, the degree of long-range order is inferior compared to the samples obtained from sedimentation. Our observations indicate that gradual sedimentation, compared to solvent evaporation, provides a driving force that is gentle and homogeneous. While sedimentation-driven assembly is not new, this is the first time that it has been used to make large-scale assemblies of polyhedral particles. In other approaches, anisotropic forces between particles primarily determine

the structure of assembled patterns [10]. The key factor in our experiments is particle shape—a feature we have found easier to control. Depletion attractions are a secondary bias that can be easily tuned by adjusting the depleting agents’ size and concentration. For octahedra this changes the packing from the Minkowski lattice to a less dense structure with more face-to-face contact. Simulations of attractive cuboctahedra also indicate that depletion attractions may be involved in the formation of the FCC-like structure observed experimentally (Supplementary Figure 3). Other polyhedral shapes that pack densely with limited face-to-face contact may be similarly prone to switching structures under the addition of excess polymer. Polyhedral nanocrystals can be made from a variety of materials, including metals, dielectrics and semiconductors [30]. Compared to crystal structures of nearly spherical particles, dense packings of polyhedra are characterized by higher packing fractions, large interfaces between particles and different geometries of voids and gaps, which determine the electrical and optical properties of these materials [31, 32]. The assembly procedures described here may thus give access to a wide range of interesting, scalable nanostructured materials with dimensions that are comparable to bulk materials.

Methods Summary

Ag polyhedra were synthesized using methods described by Tao and coworkers [18, 7] and were exceptionally monodisperse in both size and shape. The particles were suspended in DMF and loaded on top of a PDMS reservoir that had been filled with neat DMF. The entire setup was placed in a weighing bottle to minimize solvent evaporation, and particles settled to the bottom of the PDMS chamber by gravity, assembling into nanocrystal superlattices. Measurements of hydrodynamic diameter and zeta potential were obtained using commercial light scattering instrumentation (Zetasizer Nano,

Malvern Instruments). The van der Waals potential was calculated from a coarse-grained atomistic representation of the octahedra, with 666700 interaction sites per particle. For the calculation of the repulsion between polymer-covered surfaces, we assumed a dense packing of 100 spherical polymer coils per surface. Every adsorbed coil interacts with its counterpart on the opposing surface via a Gaussian core potential [33]. Effective pair potentials $w_\varphi(x)$ for octahedra in the presence of excess PVP were calculated according to $\beta w_\varphi(x) = -\ln P_\varphi(x)/P_0(x)$. Here, $P_\varphi(x)$ is the distribution function of the distance x between nearest faces of octahedra in a bath of hard spheres with packing fraction φ , and $\beta = 1/k_B T$.

References

- [1] Wells, A. F. *Structural Inorganic Chemistry* (Oxford University Press, 1984).
- [2] O’Keeffe, M. & Hyde, B. G. *Crystal Structures I. Patterns and Symmetry* (Mineralogical Society of America, 1996).
- [3] Murray, C. B., Kagan, C. R. & Bawendi, M. G. Synthesis and characterization of monodisperse nanocrystals and close-packed nanocrystal assemblies. *Annual Review of Materials Science* **30**, 545–610 (2000).
- [4] Betke, U. Densest lattice packings of 3-polytopes. *Computational Geometry* **16**, 157–186 (2000).
- [5] Haji-Akbari, A. *et al.* Disordered, quasicrystalline and crystalline phases of densely packed tetrahedra. *Nature* **462**, 773–7 (2009).
- [6] Torquato, S. & Jiao, Y. Dense packings of the Platonic and Archimedean solids. *Nature* **460**, 876–9 (2009).
- [7] Tao, A. R., Ceperley, D. P., Sinsermsuksakul, P., Neureuther, A. R. & Yang, P. Self-organized silver nanoparticles for three-dimensional plasmonic crystals. *Nano Letters* **8**, 4033–8 (2008).
- [8] Xie, S. *et al.* Supercrystals from Crystallization of Octahedral MnO Nanocrystals. *The Journal of Physical Chemistry C* **113**, 19107–19111 (2009).
- [9] Quan, Z. & Fang, J. Superlattices with non-spherical building blocks. *Nano Today* **5**, 390–411 (2010).

- [10] Glotzer, S. C. & Solomon, M. S. Anisotropy of building blocks and their assembly into complex structures. *Nature Materials* **6**, 557 (2007).
- [11] Auer, S. & Frenkel, D. Suppression of crystal nucleation in polydisperse colloids due to increase of the surface free energy. *Nature* **413**, 711–3 (2001).
- [12] Bishop, K. J. M., Wilmer, C. E., Soh, S. & Grzybowski, B. A. Nanoscale forces and their uses in self-assembly. *Small* **5**, 1600–30 (2009).
- [13] Minkowski, H. Dichteste gitterförmige Lagerung kongruenter Körper. *Nachr. Akad. Wiss. Göttingen Math.-Phys. Kl. II* 311–355 (1904).
- [14] Kabashin, a. V. *et al.* Plasmonic nanorod metamaterials for biosensing. *Nature Materials* **8**, 867–71 (2009).
- [15] Novo, C., Funston, A. M. & Mulvaney, P. Direct observation of chemical reactions on single gold nanocrystals using surface plasmon spectroscopy. *Nature Nanotechnology* **3**, 598–602 (2008).
- [16] Pendry, J. B., Schurig, D. & Smith, D. R. Controlling electromagnetic fields. *Science* **312**, 1780–2 (2006).
- [17] Hung, W. H., Aykol, M., Valley, D., Hou, W. & Cronin, S. B. Plasmon resonant enhancement of carbon monoxide catalysis. *Nano Letters* **10**, 1314–8 (2010).
- [18] Tao, A., Sinsermsuksakul, P. & Yang, P. Polyhedral silver nanocrystals with distinct scattering signatures. *Angewandte Chemie (International ed. in English)* **45**, 4597–601 (2006).
- [19] Sun, Y. & Xia, Y. Shape-controlled synthesis of gold and silver nanoparticles. *Science* **298**, 2176–9 (2002).

- [20] Chen, E. R., Engel, M. & Glotzer, S. C. Dense Crystalline Dimer Packings of Regular Tetrahedra. *Discrete & Computational Geometry* **44**, 253–280 (2010).
- [21] Kallus, Y., Elser, V. & Gravel, S. Dense Periodic Packings of Tetrahedra with Small Repeating Units. *Discrete & Computational Geometry* **44**, 245–252 (2010).
- [22] Torquato, S. & Jiao, Y. Exact constructions of a family of dense periodic packings of tetrahedra. *Physical Review E* **81** (2010).
- [23] Weeks, J. D., Chandler, D. & Andersen, H. C. Role of Repulsive Forces in Determining the Equilibrium Structure of Simple Liquids. *J. Chem. Phys.* **54**, 5237 (1971).
- [24] Ruths, M., Yoshizawa, H., Fetters, L. J. & Israelachvili, J. N. Depletion Attraction versus Steric Repulsion in a System of Weakly Adsorbing Polymers -Effects of Concentration and Adsorption Conditions. *Macromolecules* **29**, 7193–7203 (1996).
- [25] Asakura, S. & Oosawa, F. Interaction between particles suspended in solutions of macromolecules. *Journal of Polymer Science* **33**, 183–192 (1958).
- [26] Baranov, D. *et al.* Assembly of Colloidal Semiconductor Nanorods in Solution by Depletion Attraction. *Nano Lett.* **10**, 743–749 (2010).
- [27] Dogic, Z. & Fraden, S. Development of model colloidal liquid crystals and the kinetics of the isotropic-smectic transition. *Philosophical Transactions of the Royal Society A: Mathematical, Physical and Engineering Sciences* **359**, 997–1015 (2001).
- [28] Hanfland, M., Syassen, K., Christensen, N. E. & Novikov, D. L. New high-pressure phases of lithium. *Nature* **408**, 174–8 (2000).

- [29] Rabani, E., Reichman, D. R., Geissler, P. L. & Brus, L. E. Drying-mediated self-assembly of nanoparticles. *Nature* **426**, 271–274 (2003).
- [30] Ozin, G. A., Arsenault, A. C. & Cardemartiri, L. *Nanochemistry: A chemical approach to nanomaterials* (The Royal Society of Chemistry, 2009).
- [31] Talapin, D. V., Lee, J.-S., Kovalenko, M. V. & Shevchenko, E. V. Prospects of colloidal nanocrystals for electronic and optoelectronic applications. *Chemical Reviews* **110**, 389–458 (2010).
- [32] Stebe, K. J., Lewandowski, E. & Ghosh, M. Materials science. Oriented assembly of metamaterials. *Science* **325**, 159–60 (2009).
- [33] Louis, A. A., Bolhuis, P. G., Hansen, J. P. & Meijer, E. J. Can Polymer Coils Be Modeled as "Soft Colloids"? *Phys. Rev. Lett.* **85**, 2522 (2000).

Acknowledgements

This work was partially supported by DARPA. The shape selective synthesis part of the work is supported by the Director, Office of Basic Energy Sciences, Materials Sciences and Engineering Division, of the U.S. Department of Energy under Contract No. DE-AC02-05CH11231. MG and PLG were supported by the National Science Foundation under Grant No. CHE-0910981. AWC was supported by the U.S. Department of Energy under Contract No. DE-AC02-05CH11231. Work at the Molecular Foundry was supported by the Office of Science, Office of Basic Energy Sciences, of the U.S. Department of Energy under Contract No. DE-AC02-05CH11231. P. Y. thanks NSF for the Waterman Award.

Author contributions

JH and MG both contributed extensively to this work. JH initiated the study, conceived and conducted the experiments, analyzed results and co-wrote the paper. MG identified all crystal structures, conceived and implemented the simulations and theoretical models, analyzed results and co-wrote the paper. AWC suggested simulations and experiments, analyzed results, helped prepare figures, and cowrote the paper. PLG suggested simulations and experiments, analyzed results and co-wrote the paper. PY initiated the study, suggested experiments, analyzed results and co-wrote the paper.

Additional Information

The authors declare no competing financial interests. Supplementary information accompanies this paper.

Figure legends

Figure 1. Self-assembly of dense polyhedra lattices. **a**, Scheme showing polyhedral shapes accessible using the Ag polyol synthesis. **b**, Microfluidic reservoir used to assemble the Ag nanocrystals. **c**, Dark field micrograph shows the growing Ag supercrystal. **d-h**, SEM micrographs of the colloidal lattices (left) and their corresponding diagrams of the densest known lattice packings (right): **d**, cubes, **e**, truncated cubes, **f**, cuboctahedra, **g**, truncated octahedra, **h**, octahedra. Scale bars are 500 nm unless otherwise noted.

Figure 2. Long-range order of assembled lattices. Large-scale SEM images of an octahedra lattice on a Si substrate. **a**, The imaged region is a cross-section of a large supercrystal that extends millimeters below the plane of the image, and was exposed by cleaving the substrate and superlattice into two pieces. **b-d**, Inspection of particle orientation in the magnified sections reveals that long-range order is maintained across the entire imaged sample. **Inset b-d**, Fourier transforms of these sections match closely. **e**, SEM image of a large Minkowski facet. **f**, A magnified section of the experimental lattice with a superimposed rendering of the ideal Minkowski lattice shows excellent agreement. **g**, Cross-polarized optical micrograph of a millimeter-sized superlattice slab. The uniform color of the reflected light from top to bottom indicates that the crystal is composed of a single domain. All scale bars are 4 microns unless otherwise noted.

Figure 3. Gravitational driving force and packings near surfaces. **a**, Snapshot of a Monte Carlo simulation (Supplementary Methods) of hard octahedra under the influence of gravity (left). Zoom-in images (right), from top to bottom, show a dilute fluid, a dense fluid, and a growing solid of the Minkowski type. **b**, Packing fraction as a function of Monte Carlo sweeps and height along the z -axis in the elongated box. A density gradient develops in the initially homogenous system, resulting in a fluid-solid transformation at

the bottom. Biased by a flat surface, octahedra form a simple hexagonal lattice (space group P6/mmm). **c**, Small patches of this structure, retrieved early in the assembly process. **d**, Top-down view (left) with inset showing a vacancy in the lattice, and side view (right). **e**, Zero gravity, isobaric Monte Carlo simulations with a hard bottom wall and weak interparticle attractions show coexistence between the layered structure and the Minkowski lattice, consistent with experiment. **f**, Comparison between a flat layer of the simple hexagonal lattice (left) and corresponding staggered layers in the Minkowski lattice (right). All scale bars are 1 μm .

Figure 4. Polymer-mediated nanocrystal interactions. **a**, Contributions to the free energy of two octahedra aligned face-to-face, as a function of the separation X of surface adsorbed polymer layers (see inset) (Supplementary Methods). The entropic repulsion (orange) of densely packed surface polymers (20 nm diameter) dominates the van der Waals attraction (red) of the silver cores. The resulting potential (green) is only weakly attractive, allowing the nanocrystals to behave as quasi-hard particles. Unadsorbed polymer coils act as depleting agents, inducing effective attractions (blue) that can lead to self-assembly of different structures. **b**, Potentials of mean force for two hard octahedra in a bath of hard 20 nm spheres (see inset), as a function of the distance x between nearest midpoints of octahedra faces (see inset, interaction centers are marked yellow), averaged over all rotational degrees of freedom. Legend values correspond to the sphere packing fraction φ .

Figure 5. Ag octahedra assemble into a novel lattice in the presence of excess PVP. **a**, SEM micrograph of a new octahedra supercrystal. **b**, The lattice consists of tetramer motifs and, **c**, forms spontaneously in Monte Carlo simulations of octahedra with depletion attractions. **d**, The tetramers can be decomposed into lines and counter-

rotating helices, where the helices in this plane consist of particles belonging to 3 other sets of interlocking lines. **e**, The lattice has been colored to show these 4 equivalent sets of lines (yellow, blue, green, and red) that are oriented in different $[111]$ directions. Lines belonging to the same $[111]$ direction are parallel and shifted relative to one another in alternating layers (green particles on left; right zoom-in SEM micrograph from the corresponding experimental lattice). When viewed end-on (yellow particles), the particle lines stack hexagonally.

Acknowledgements: This work was supported by the Director, Office of Science, Office of Basic Energy Sciences, Material Sciences and Engineering Division, of the U.S. Department of Energy under Contract No. DE-AC02-05CH11231. We thank the National Center for Electron Microscopy for the use of their facilities.

Figure 1

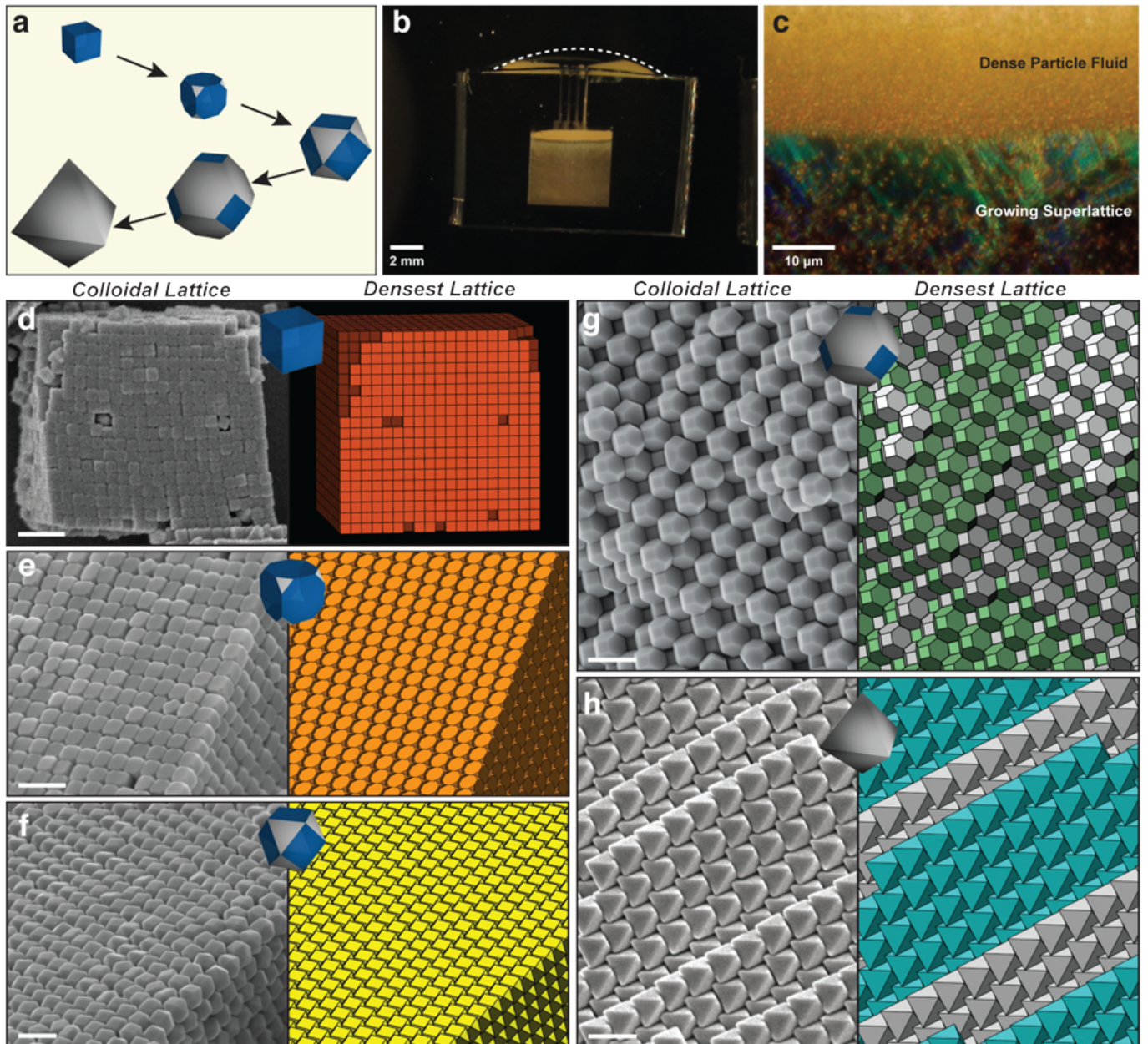


Figure 2

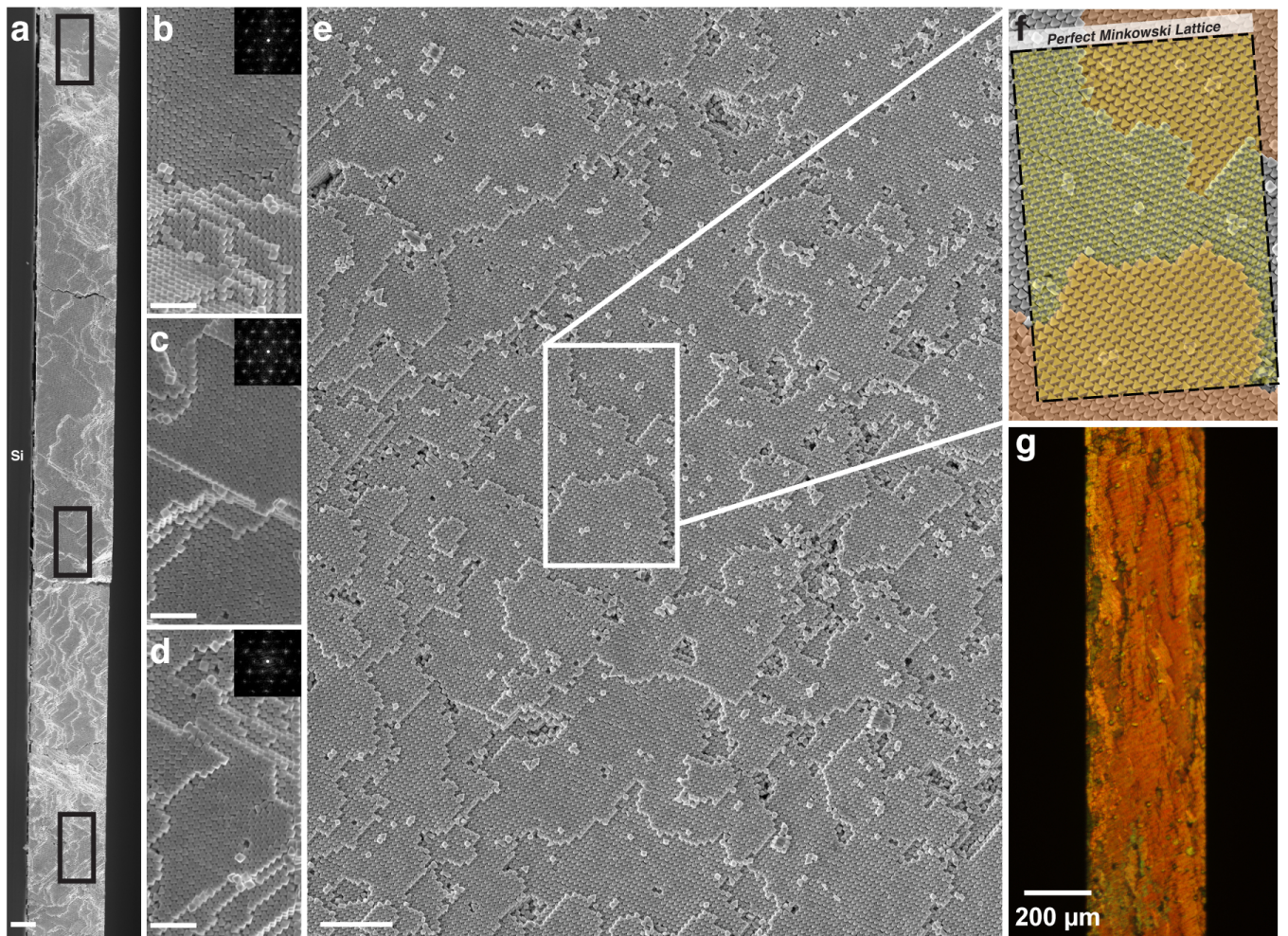


Figure 3

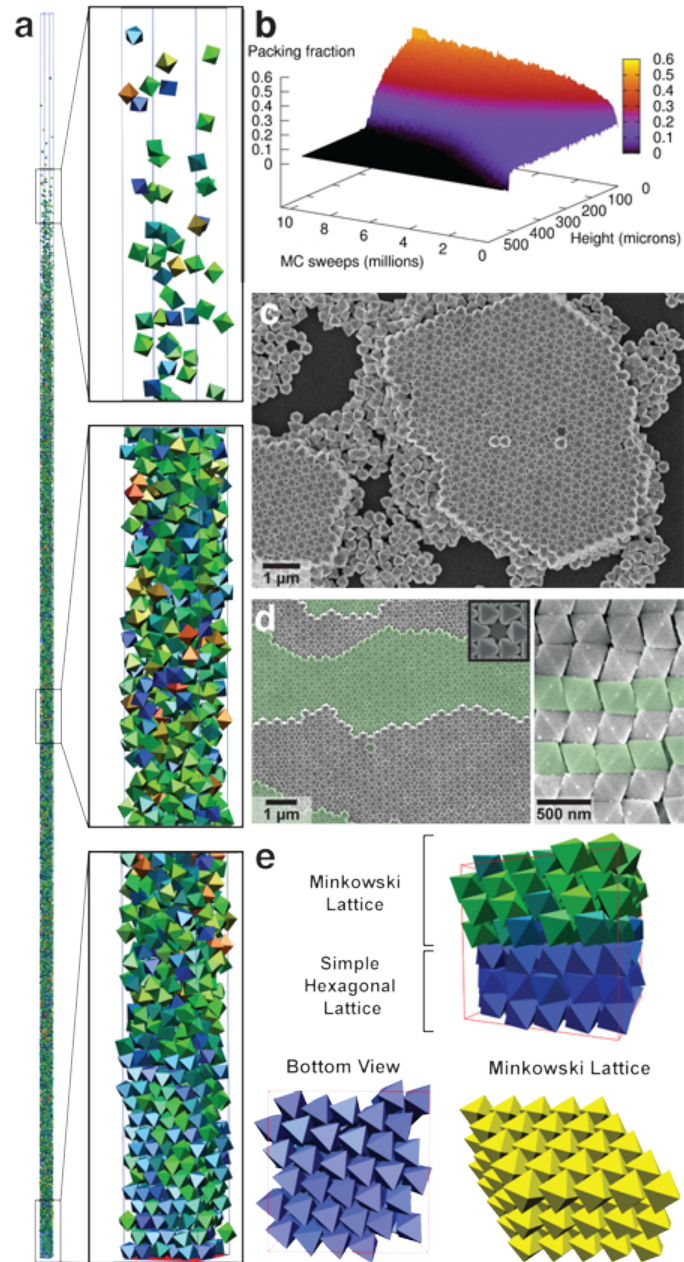


Figure 4

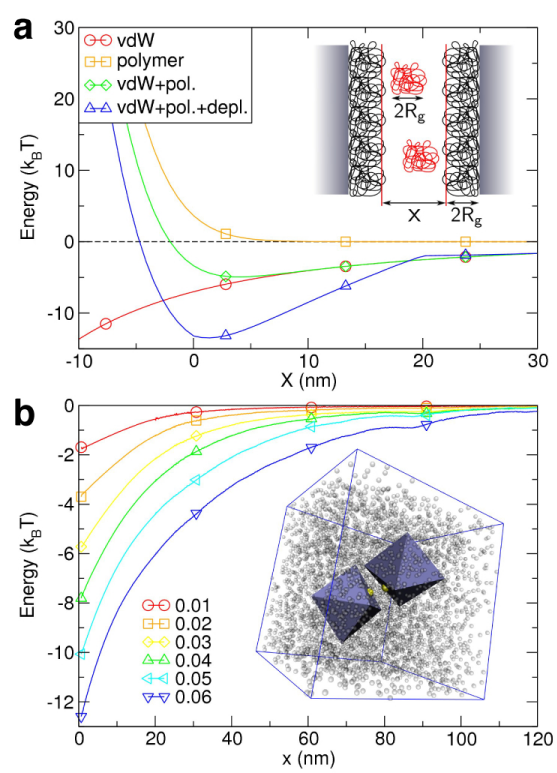


Figure 5

